**AIM:** To predict and test a range of precipitation reactions

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| **MATERIALS:**   * M solutions of magnesium nitrate, potassium iodide, silver nitrate, sodium hydroxide and iron sulfate in dropper bottles * Gloves * Safety glasses * Laminated precipitation test sheet * Paper towel | **METHOD:**  *Wear safety glasses and follow any other safety precautions given by your teacher.*   1. Using the solubility chart on the back of this page, predict whether a precipitate will form from each combination of solutions. Record these predictions in the below results table. 2. For the first pair of solutions place one drop of each on the precipitation test sheet. 3. Move the mixture so that it covers both the black and white section of the test area. 4. Record your observations in the below. Did a precipitation form? If so, did it have a colour? 5. Remove solution combination from the precipitation test sheet with paper towel. 6. Repeat steps 2-6 for the remaining solution combinations. |

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| **Solution A** | **Solution B** | **Prediction: Will a precipitate form?** | **Name of precipitate (if any)** | **Observations** |
| magnesium nitrate | potassium iodide |  |  |  |
| magnesium nitrate | silver nitrate |  |  |  |
| magnesium nitrate | sodium hydroxide |  |  |  |
| magnesium nitrate | iron sulfate |  |  |  |
| potassium iodide | silver nitrate |  |  |  |
| potassium iodide | sodium hydroxide |  |  |  |
| potassium iodide | iron sulfate |  |  |  |
| silver nitrate | sodium hydroxide |  |  |  |
| silver nitrate | iron sulfate |  |  |  |
| sodium hydroxide | iron sulfate |  |  |  |

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